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Molecular weight of the gas = 29 ;

$1 = 1.4 \cdot 1.4 \cdot 7 - \dots - \text{£} 1 \cdot 2 \cdot \text{' Li} 2 \cdot 14 \cdot 1$

$P_1 \cdot [1 + 1J - [L4 - +1) : o'528 p, =$

$430 \cdot 1 : 0.523 = 253.44 \text{ kPa}$ Fig.

$11121 \cdot 1', 2 \cdot 2 \cdot T_1 \cdot 7 + 1 \cdot 1.4 + \}$

$do \cdot 7', = 303 \times 0.3333 : 573.3 \text{ K R}$

$= 9 - \text{' S; } - \text{' 3} = 0.2357 \text{ kJ/ kg} \cdot \text{C}$ _

$R_1 \cdot 0.2867 \times 1.4 \cdot cp - \text{---} 1 \cdot | -$

$\text{-----} | - 1.0034 \text{ kJ/ kg} \cdot \text{C} \cdot \text{C} \cdot 2$

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